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between molarity and
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Quiz & Worksheet -
How to Calculate
Molarity and Molality
Answers

Key. 1) 23.5g of NaCl is dissolved in enough water to make 683 L of solution. a) What is the molarity (M) of the solution? Molar mass of NaCl = 58.44 g/mole
Moles of NaCl: $\frac{23.5 \text{ g NaCl}}{58.44 \text{ g/mole}} = 1 \text{ mole NaCl}$

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Molarity And

Worksheet With
Answers

0.402 moles NaCl
58.44 g NaCl
Molarity =
moles /
= 0.402 moles NaCl /
L = 0.589 M NaCl
liters solution
0.683 L of solution
b) How many moles of NaCl are contained in 0.0100 L

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Molarity And

of the above NaCl solution? + + + 0.

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Molarity Molality

Osmolality

Osmolarity

Worksheet and Key ...

Calculate molarity of

35.0 mL KOH

solution needed to

completely neutralize

22.5 mL of 1.75 M H

2SO_4 . Calculate

volume (mL) of

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Molarity And

2.50M H₂SO₄

needed to completely
neutralize 10.0g

NaOH (s). Answers. M

$M_1 V_1 = M_2 V_2$ (1.71

M)(25.0 mL) = M₂

(65.0 mL) M₂ =

0.658 M; M = mol/L =

$(25.0/40.0) / (0.325)$

= 1.92 mol/L

Molarity 1

(Worksheet) -

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25.0g NaOH in 325
mL of solution.

Calculate grams of
solute needed to
prepare 225 mL of
0.400 M KBr solution.
Calculate mL of

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Worksheet

MOLARITY (M) = $\frac{\text{moles of solute}}{\text{L of solution}}$

MOLALITY (m or)

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Molarity And

= moles of solute

Liters of solvent kg of solvent Molarity

Example: 4.0 moles of LiCl is dissolved in 5.0 liters of water.

What is the molarity of the solution? $4.0 \text{ moles} = 0.8 \text{ M}$ This solution is 0.8 Molar or 0.8 M 5.0 Liters

7) How many moles of solute are in 125

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Molarity And

mL of a 2.0 M ...

Mole

Fraction/Molality

Worksheet. Name:

Date: 1. A solution is prepared by mixing 100.0 g of water, H_2O , and 100.0 g of ethanol, C_2H_5OH .

Determine the mole fractions of each substance. 2. The molality of an aqueous solution of

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sugar ($C_{12}H_{22}O_{11}$)

is 1.62m. Calculate the mole fractions of sugar and water.

Chemistry 11 Mole

Fraction/Molality

Worksheet Date

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2 identifiera _____

What does molarity

mean? Number of

moles of solute. 1

liter solution. What is

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Molarity
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Answers

the molarity of a solution that contains 4.53 moles of lithium nitrate in 2.85 liters of solution? $4.53 \text{ mol LiNO}_3 = 1.59 \text{ M}$
 $4.53 \text{ mol LiNO}_3 / 2.85 \text{ L soln}$

Molarity Worksheet 2
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This quiz and
worksheet allow
students to test the

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following skills:

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acquired knowledge
to answer practice
problems involving
the calculation of
molality Information
recall -...

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The density of the
solution is 0.993

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g/mL. What is the molarity, molality and mole fraction of acetone in this

solution? Solution: 1)

Preliminary

calculations: mass of acetone: (3.30 mL)

(0.789 g/mL) =

2.6037 g moles of

acetone: 2.6037 g /

58.0794 g/mol =

0.04483 mol --- need

to look up formula of

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Worksheet With ChemTeam: Molality Problems #1-10

PDF (4.32 MB) This is a single 2-page worksheet for preparing solutions, interpreting and drawing particle diagrams, and molarity calculations. There are a total of 5 questions. Answer

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key is included. The download includes a handout master (.pdf) that includes one worksheet, and answer key. This product is designed.

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problems on each worksheet using a customized algorithm to produce problems that are realistic and unique. Many aspects of the problems can be customized to best fit the needs of your students. You have the option to present the amount of solutes to the students as moles, grams, or a

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mixture of both.

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Molarity: The molarity of a solution is calculated by taking the moles of solute and dividing by the liters of solution.

Molarity is designated by a capital "M".

Molarity = Moles
Solute / Liter of

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Molarity And

Solution. Molality:

The molality of a solution is calculated by taking the moles of solute and dividing by the kilograms of solvent. Molality is designated by a lower case "m".

Molarity and Solution
Units of
Concentration

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Molarity And

Molality provides many examples for students to practice calculations involving Molarity & Molality. A complete answer key is provided at the end. This worksheet can be used in any Chemistry class, regardless of the students' ability level.

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Molality
following questions in
the space provided.

Molarity is a
temperature

dependent quantity,

whereas molality is

not. Page 9/27

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Answers

$$MV = \text{grams} / \text{molar mass. (x) (1.000 L) = 245.0 g} / 98.0768 \text{ g mol}^{-1} \cdot x =$$

2.49804235 M. to four sig figs, 2.498 M.

If the volume had been specified as 1.00 L (as it often is in problems like this), the answer would have been 2.50 M, NOT 2.5 M.

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ChemTeam: Molarity
Problems #1 - 10

Molarity, also known
as molar

concentration, is the
number of moles of a
substance per liter of
solution. Solutions
labeled with the
molar concentration
are denoted with a
capital M. A 1.0 M
solution contains 1

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Molality
mole of solute per
liter of solution.

Molality is the
number of moles of
solute per kilogram of
solvent.

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