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Gases

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11 REVIEW Gases SECTION 3 SHORT  
ANSWER Answer the following questions  
in the space provided. 1. \_\_\_\_\_ The molar  
mass of a gas at STP is the density of that  
gas (a) multiplied by the mass of 1 mol. (c)  
multiplied by 22.4 L. (b) divided by the  
mass of 1 mol. (d) divided by 22.4 L. 2.

Chapter 11 Review Gases Section 3  
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Three of the primary components of air are carbon dioxide, nitrogen, and oxygen. In a sample containing a mixture of only these gases at exactly 1 atm, the partial pressures of carbon dioxide and nitrogen are given as  $P_{\text{CO}_2} = 0.285$  torr and  $P_{\text{N}_2} = 593.525$  torr. What is the partial pressure of oxygen?



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SECTION 1 continued 5. Convert a  
pressure of 0.200 atm to the following  
units: 152 a. mm Hg 20.3 b. kPa 6. When  
an explosive like TNT is detonated, a  
mixture of gases at high temperature is  
created. Suppose that gas X has a pressure

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of 50 atm, gas Y has a pressure of 20 atm, and gas Z has a pressure of 10 atm. 80 atm  
a. What is the total ...

mc06se cFMsr i-vi

Modern Chemistry 99 Gases CHAPTER  
11 REVIEW Gases SECTION 4 SHORT  
ANSWER Answer the following questions  
in the space provided. 1. List the following  
gases in order of increasing rate of  
effusion.

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CHAPTER 11 REVIEW Gases  
SECTION 3 SHORT ANSWER Answer  
the following questions in the space  
provided. 1. \_\_\_\_\_ The molar mass of a  
gas at STP is the density of that gas (a)  
multiplied by the mass of 1 mol. (c)

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multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided ...

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the space provided, write the letter of the  
term or phrase that best completes each  
sentence or best answers each question.

\_\_\_\_\_ 1. At the same temperature and  
pressure, balloons of equal volume always  
contain

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