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Chapter 10 Supplemental Problems Balance the following chemical equations. 1) $\text{SnS}_2 (\text{s}) + \text{O}_2 (\text{g}) \rightarrow \text{SnO}_2 (\text{s}) + \text{SO}_2 (\text{g})$ 2) $\text{C}_2\text{H}_6 (\text{g}) + \text{O}_2 (\text{g}) \rightarrow \text{CO}_2 (\text{g}) + \text{H}_2\text{O} (\text{g})$ 3) $\text{Al} (\text{s}) + \text{HCl} (\text{aq}) \rightarrow \text{AlCl}_3 (\text{aq}) + \text{H}_2 (\text{g})$ 4) $\text{CoCO}_3 (\text{s}) \rightarrow \text{CoO} (\text{s}) + \text{CO}_2 (\text{g})$ Write a balanced equation for each of the following reactions, substituting symbols and formulas for names.

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CHAPTER 10 SOLUTIONS MANUAL 3. Apply How many atoms of hydrogen-1 are in a 1.007-g sample? Recall that 1.007 amu is the mass of one atom of hydrogen-1. Round your answer to two significant digits. 1.007 g 1 H-1 atom ___ 1 ...

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